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15 Acid-Base Titration and pH

Modern Chemistry Chapter 15 Test Review 25 multiple choice definition of self-ionization  $[H_3O^+][OH^-] = 1 \times 10^{-14}$  definition and formula for calculating  $pH = -\log[H_3O^+]$   $pH > 7$  is a base;  $pH < 7$  is an acid;  $pH = 7$  is neutral pH range is normally 0 to 14 calculate pH from  $[H_3O^+]$  calculate  $[H_3O^+]$  from pH definitions of indicators, transition interval, titration, neutralization, endpoint, standard ...

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Which of the following masses is the largest? A. 0.200 g ...

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